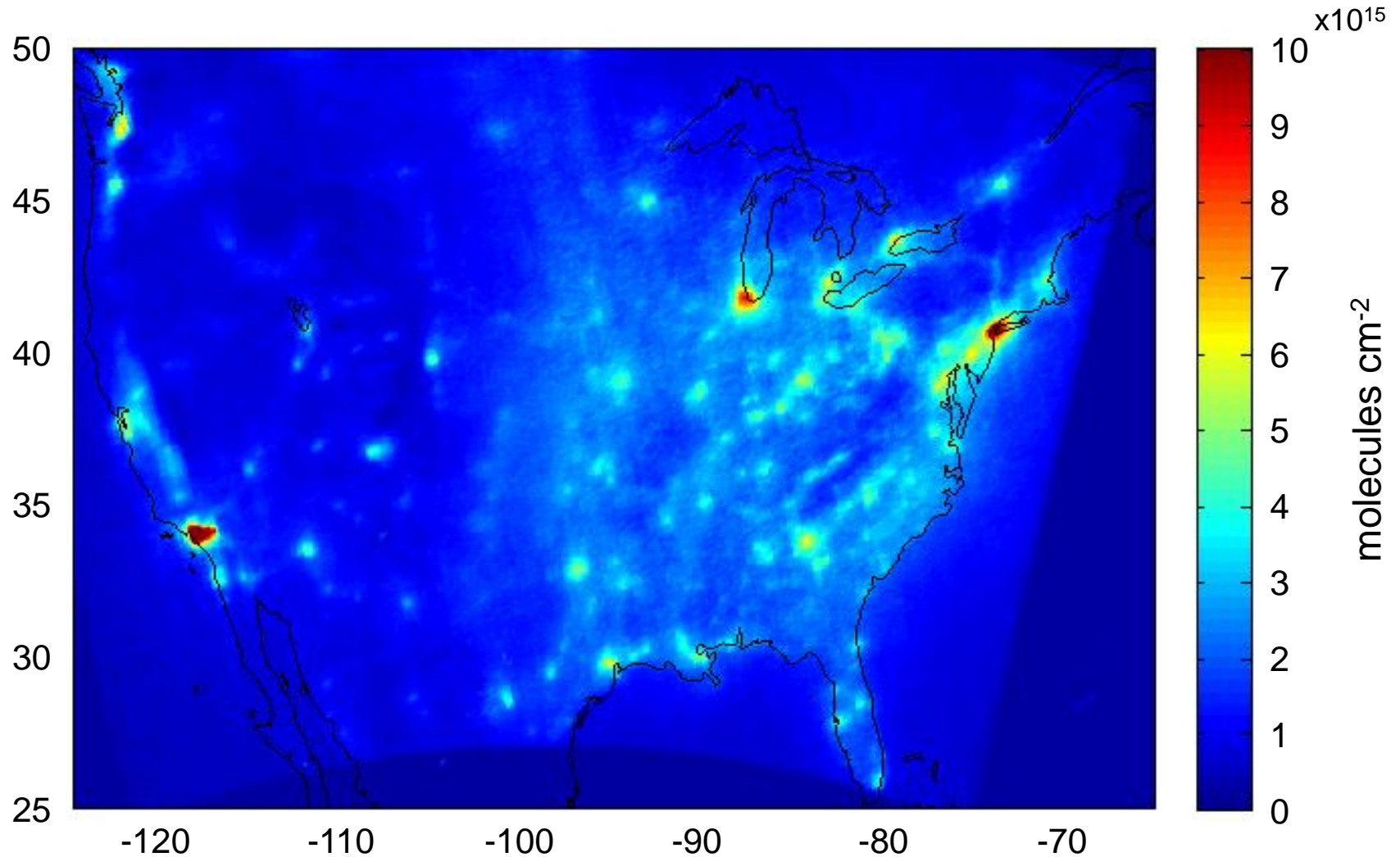


# **NO<sub>x</sub> Chemistry during WINTER**

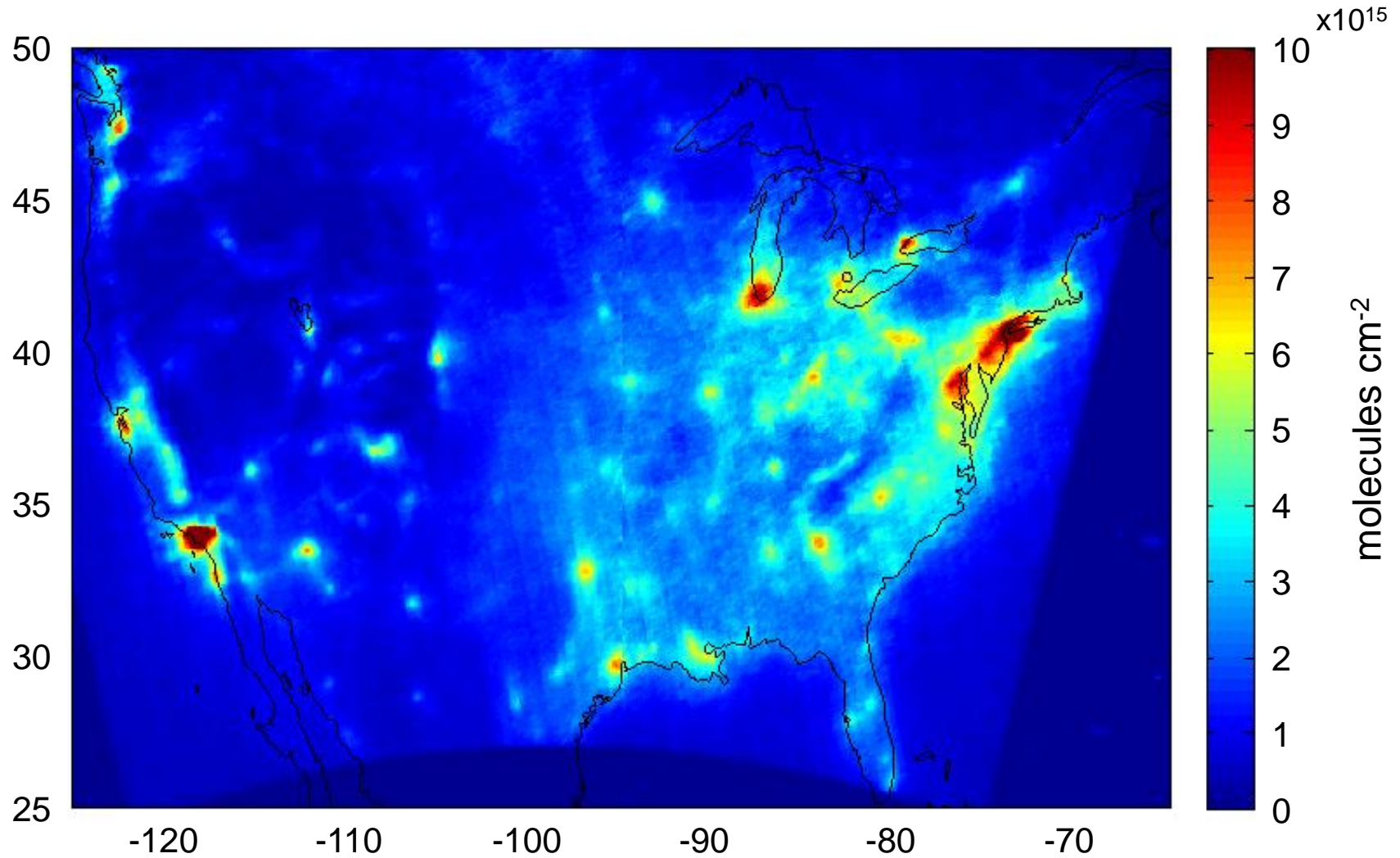


**Carlana Ebben  
Prof. Ron Cohen  
UC Berkeley**

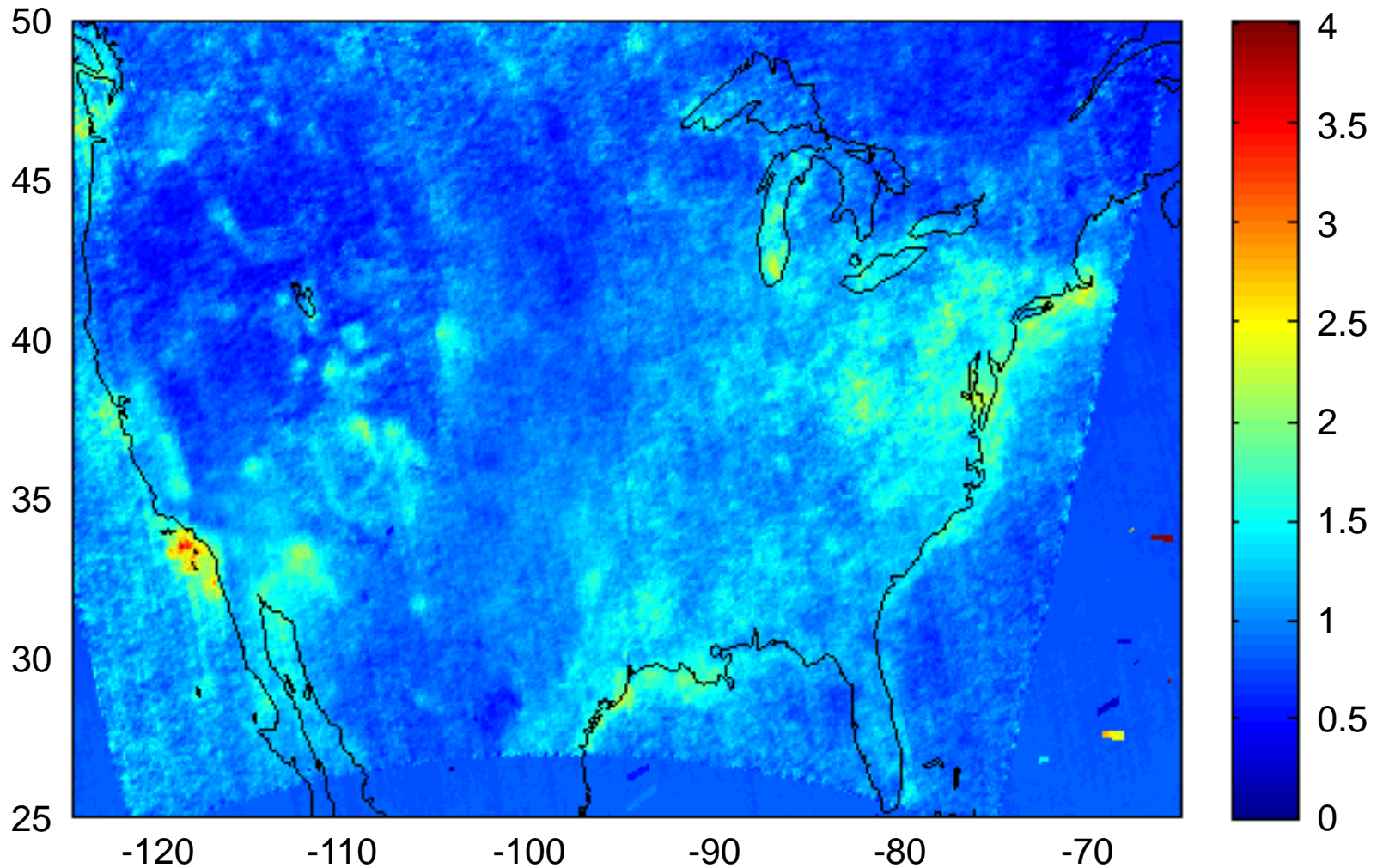
# Summer-average NO<sub>2</sub> columns in 2011



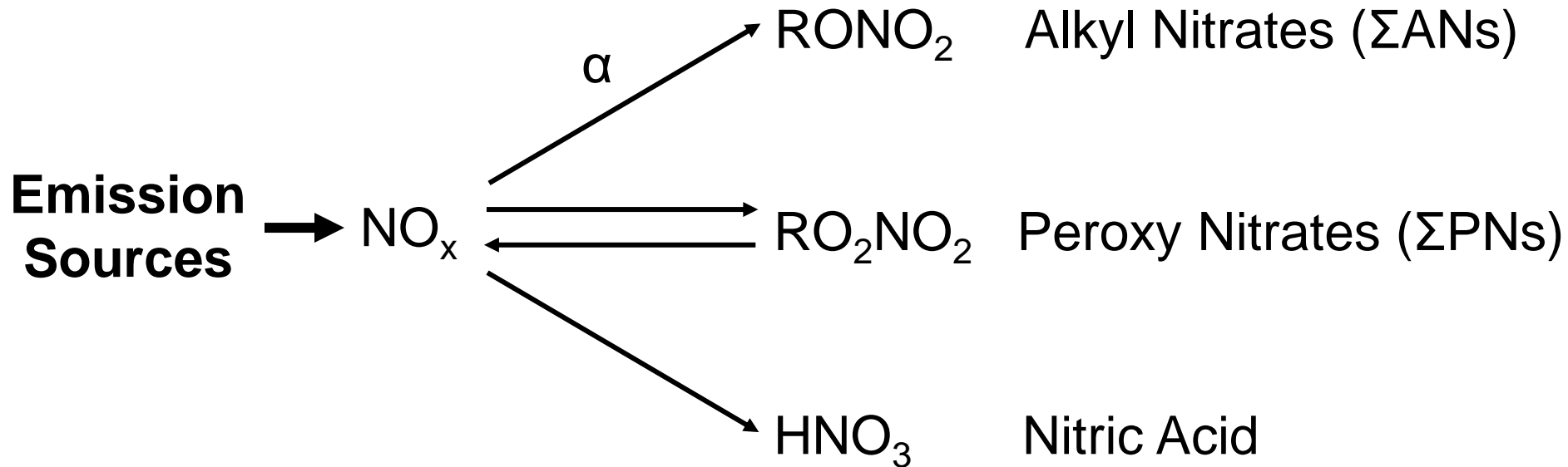
# Winter-average $\text{NO}_2$ columns in 2011



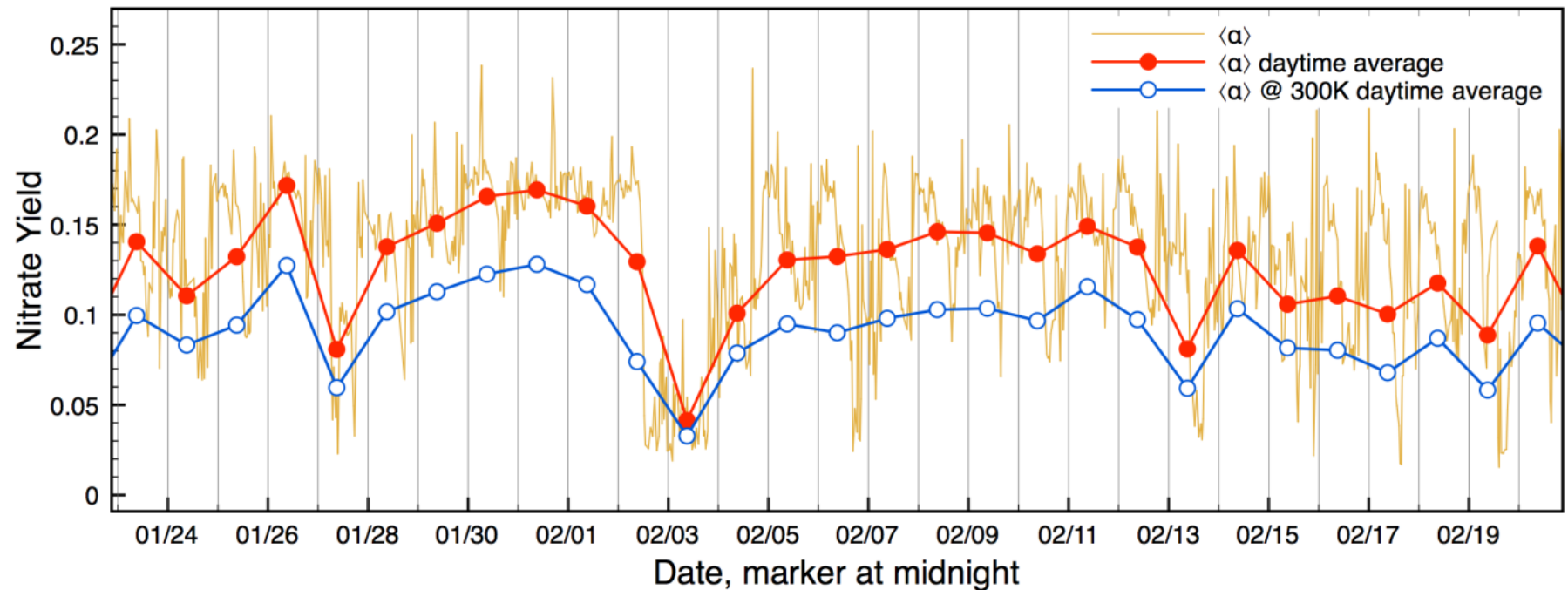
# The ratio of winter to summer $\text{NO}_2$ columns in 2011



# $\text{NO}_x$ lifetime is controlled by its sinks

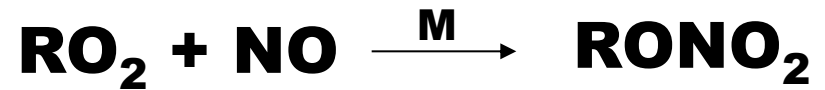
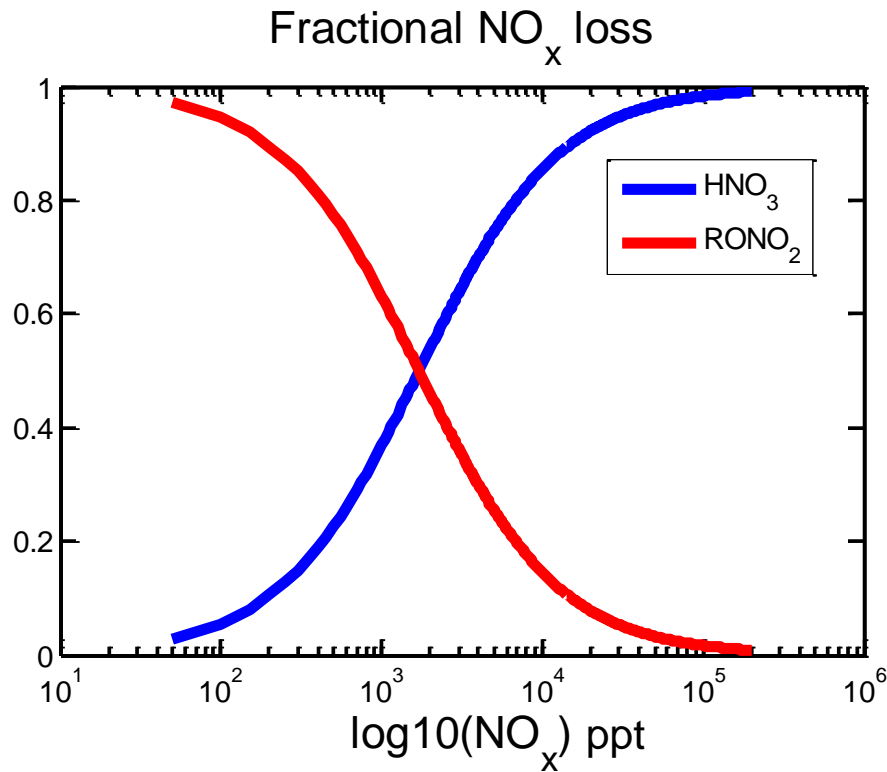


# Temperature dependence of alkyl nitrate formation branching ratio ( $\alpha$ )

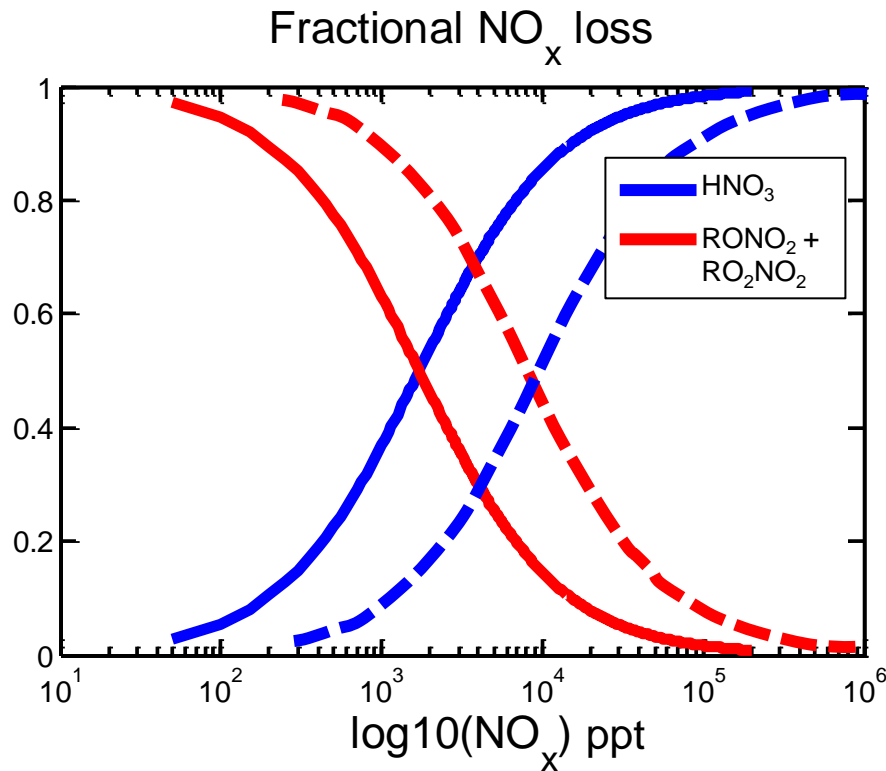


Utah Basin, Winter 2012

# Sinks of $\text{NO}_x$



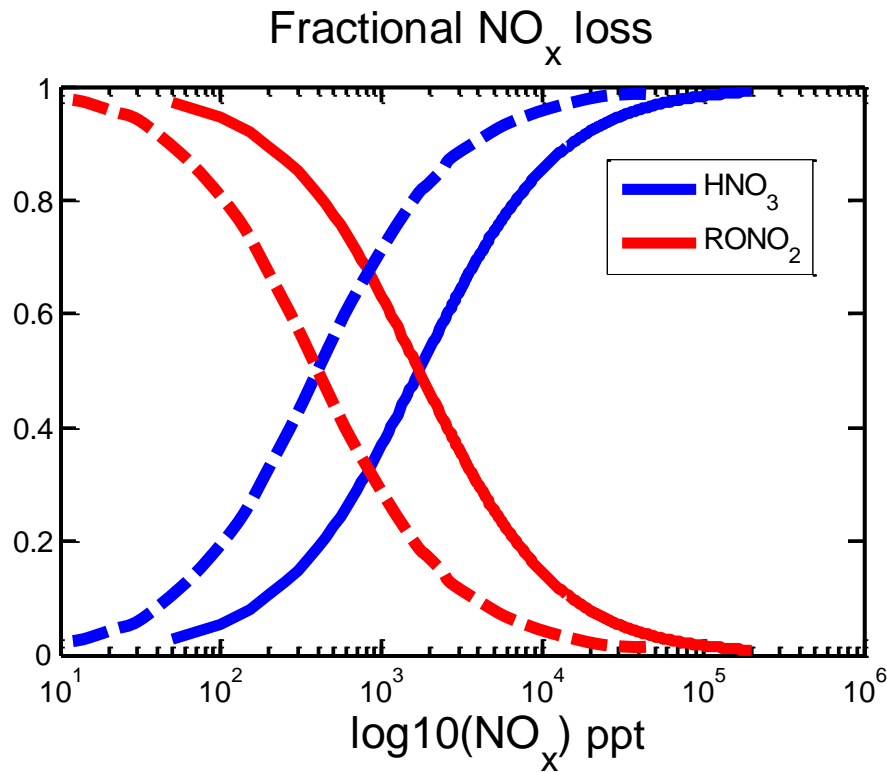
# Sinks of NO<sub>x</sub>



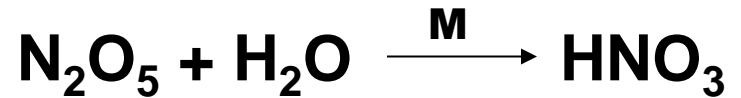
In winter, peroxy nitrate lifetimes are longer.



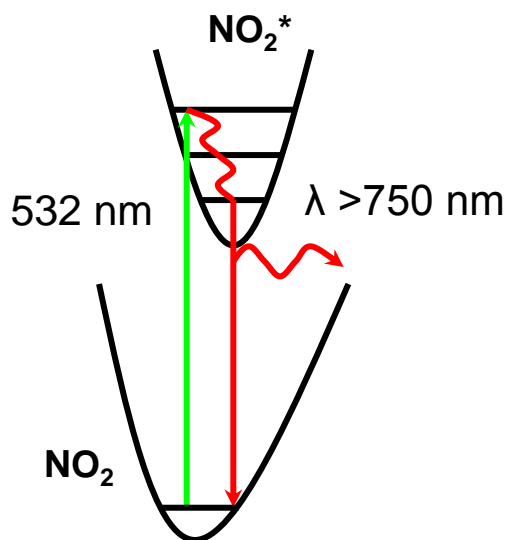
# Sinks of $\text{NO}_x$



At night:

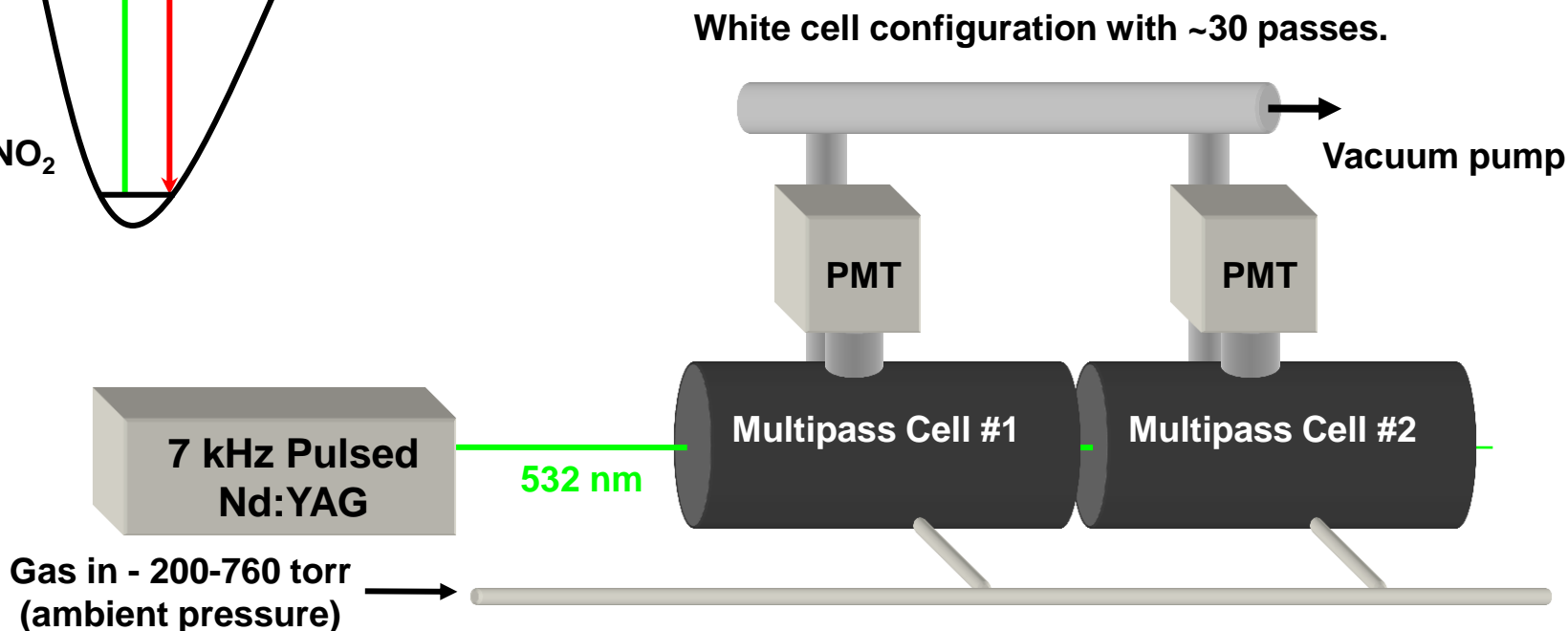


# Laser-induced fluorescence (LIF) detection of NO<sub>2</sub>



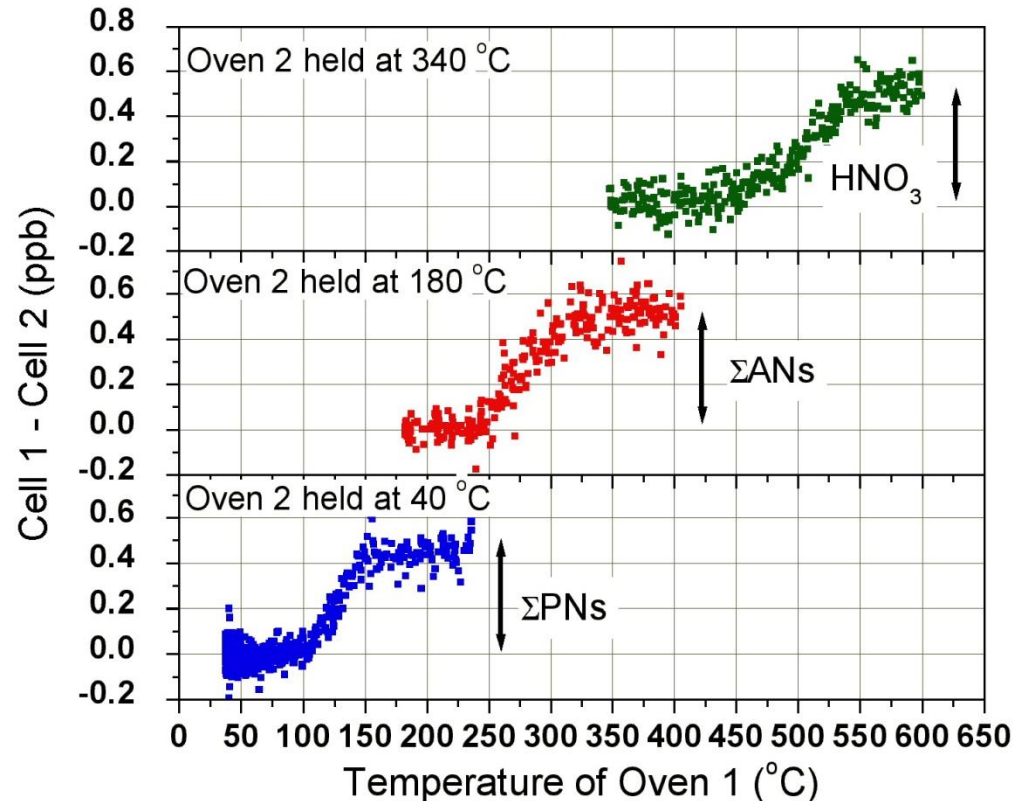
The limit of detection of the instrument is 25 pptv/10 s for NO<sub>2</sub> at S/N=2.

- Calibration done using NO<sub>2</sub> standard
- Zero air used to obtain instrument zero

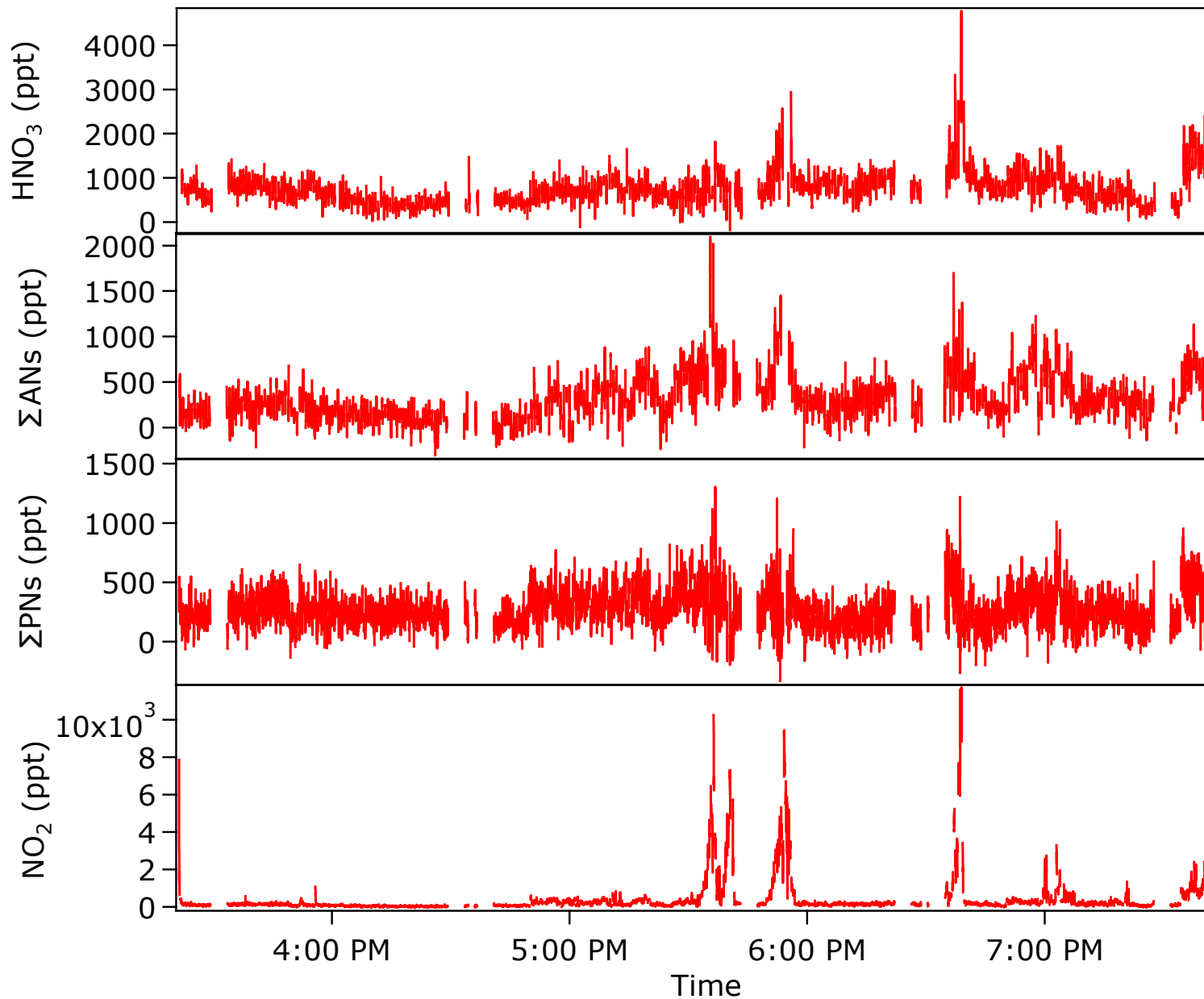


# Thermal dissociation laser-induced fluorescence (TD-LIF)

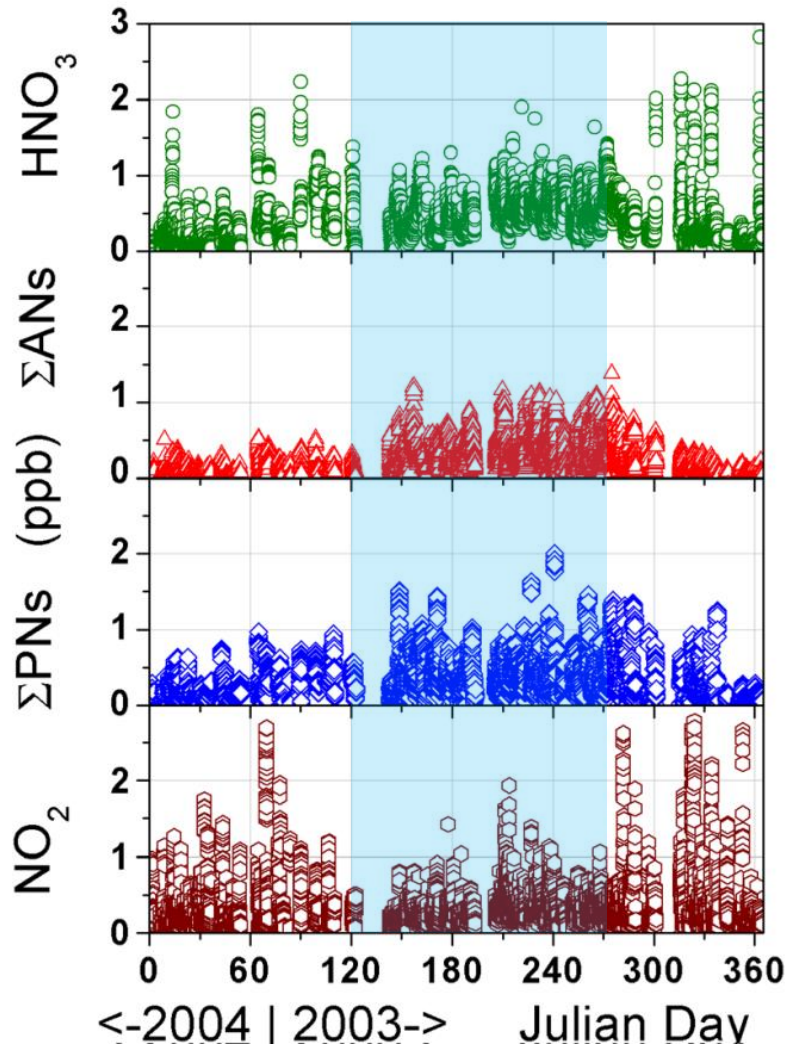
- $XNO_2 + \text{heat} \rightarrow \boxed{NO_2} + X$
- Differing bond strengths lead to dissociation at characteristic temperatures.
- Peroxy nitrates ( $RO_2NO_2$ ) and  $N_2O_5$  dissociate at  $\sim 180^\circ\text{C}$ .
- Alkyl nitrates ( $RONO_2$ ) and  $ClNO_2$  dissociate at  $\sim 340^\circ\text{C}$ .
- Nitric acid ( $HNO_3$ ) dissociates at  $\sim 600^\circ\text{C}$ .



# TD-LIF data

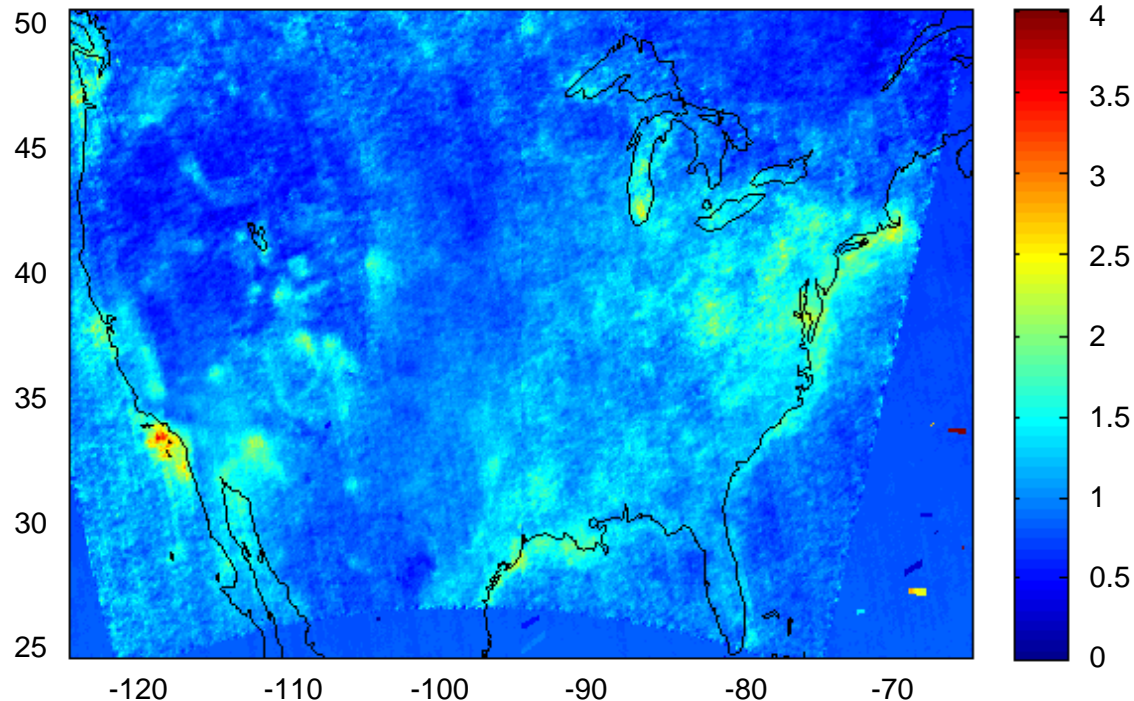


# Seasonable variability of N chemistry



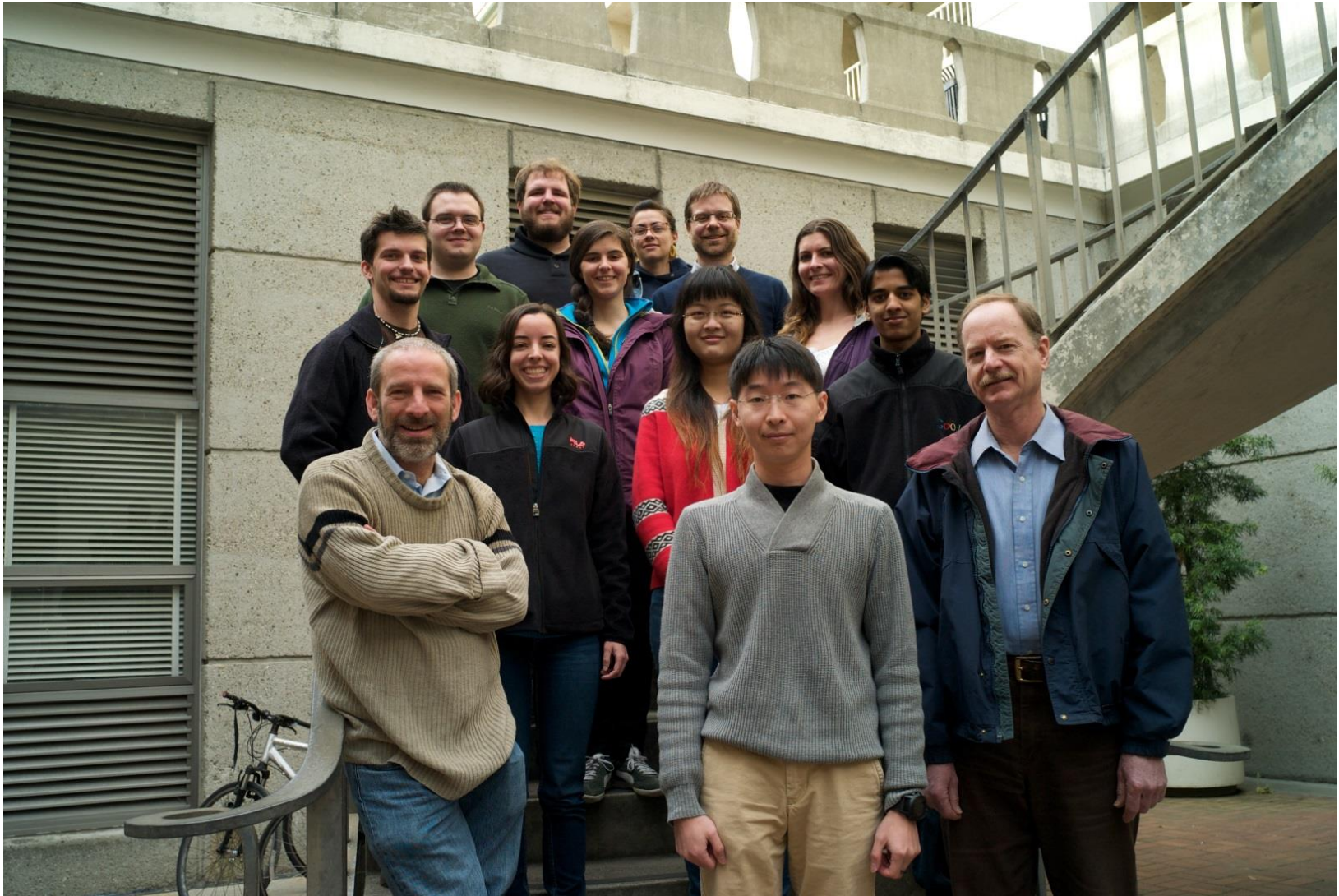
Sierra Nevada  
Mountains,  
2003-2004

# Research Questions



How are  $\text{NO}_x$  emissions, transport, and chemistry influenced by differences in photochemistry, oxidation, and temperature during the winter?

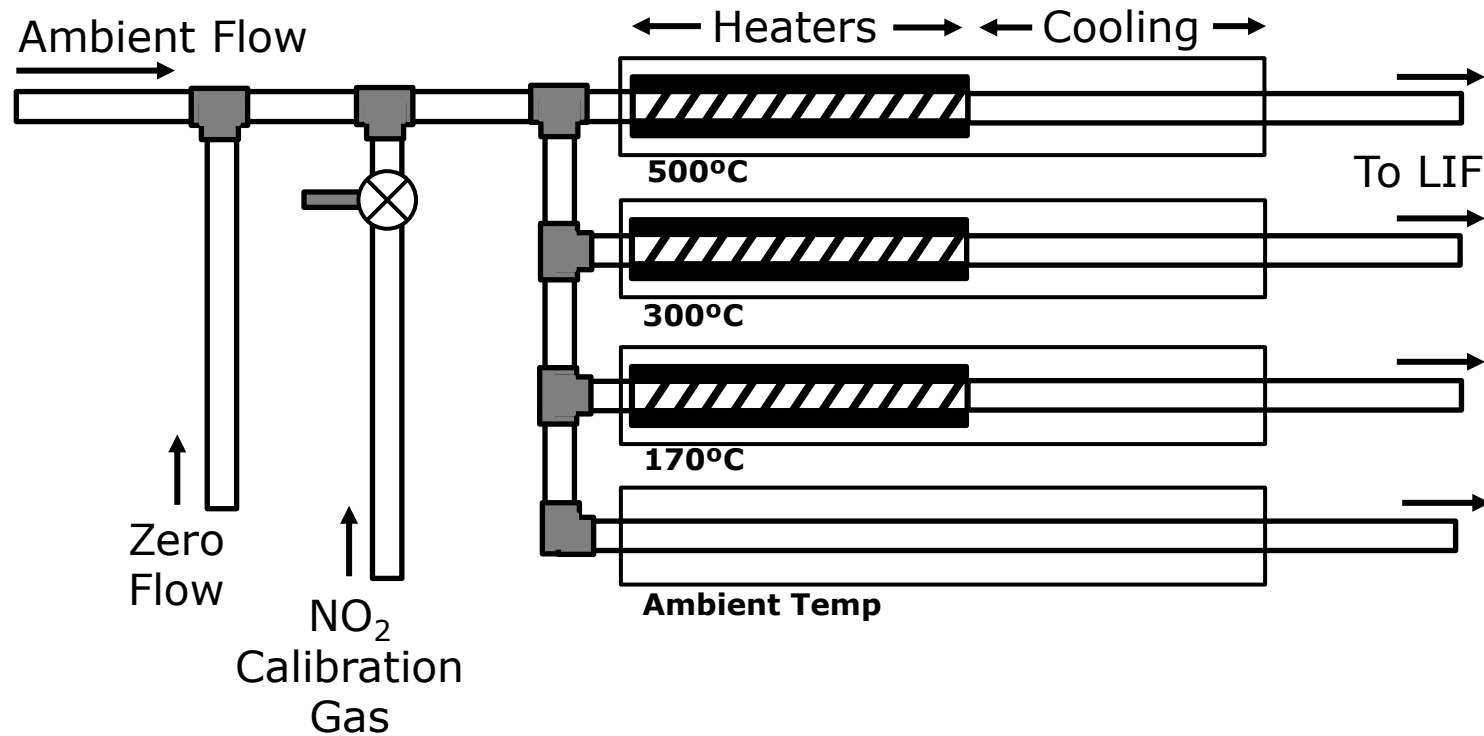
# Thank you for your attention!







# TD-LIF detection of nitrogen oxides



- Calibration done using NO<sub>2</sub> standard
- Zero air used to obtain instrument zero